Cycling battery with reference electrode using the PAT-cell test cell

I – INTRODUCTION

Until recently, to study both the positive and the negative electrode of the battery, researchers investigated half-cells. Now it is becoming more and more common to study a battery with a reference electrode. [1-2] With this configuration, researchers can get information simultaneously from both electrodes. To achieve this goal, researchers need advanced testing cells along with advanced potentiostats/galvanostats. Because of the outstanding reliability of the built-in lithium metal reference electrode, the PAT-Cell is the ideal test cell for long-term 3-electrode experiments on Li-ion battery syst-ems. In this device, the user can build an experiment cell and test the materials of the cathode and anode electrode. The Bio-Logic potentio-stat/galvanostat is a perfect match for controlling this type of experiment, as it is capable of monitoring the half-cell voltages while controlling the full cell voltage. In this note, to show the advantage of the PAT-Cell equipped with reference electrode, we describe a typical cycling experiment of a lithium-ion battery comprised of an NMC (Nickel Manganese Cobalt) cathode and a graphite anode (both with a capacity of ~2 mA.h.cm^{-2}, purchased from CCI) in a conventional LiPF$_6$ based electrolyte (1 molL^{-1} in Ethylen Carbonate/ DiMethyl-Carbonate 1:1 with 2% Vinylen Carbonate, BASF).
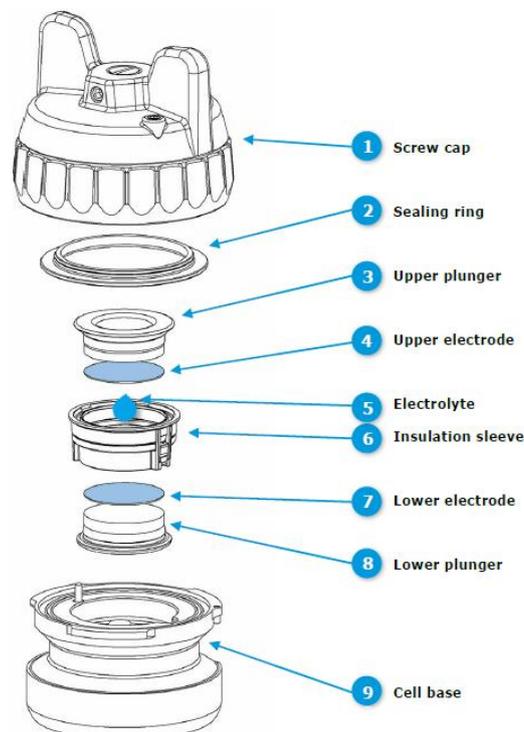
II – EXPERIMENTAL SET-UP

Fig. 1 depicts the PAT-Cell docked into the PAT-Single-Stand [4], and connected to the Bio-Logic VSP potentiostat galvanostat.

A

B

1	Screw cap
2	Sealing ring
3	Upper plunger
4	Upper electrode
5	Electrolyte
6	Insulation sleeve
7	Lower electrode
8	Lower plunger
9	Cell base

Figure 1: A) Experimental set-up with the PAT-Cell docked into the PAT-Single-Stand (in the red rectangle) connected to the VSP potentiostat. B) Exploded view of the PAT-Cell.

EC-Lab® software provides several powerful GCPL techniques which can be used for the cycling of batteries with Constant Current/ Constant Voltage (CC/CV) including sophis-ticated Galvanostatic Intermittent Titration Techniques (GITT). GCPL is the acronym for Galvanostatic Cycling with Potential Limi-tation. The GCPL techniques differ in several options such as the available step end conditions and the potential control modes used. Differences between GCPL techniques are explained in the Technical Note #30 [5].

In the present experiment, we have applied the GCPL6 technique, which allows control of the full cell voltage between the NMC cathode (socket 1 at the PAT-Single-Stand) and the graphite anode (socket 2). At the same time, the GCPL6 technique is recording the two half-cell voltages and the voltage of the full cell:

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• The voltage between NMC and the lithium reference (sockets 1S and R of the PAT-Single-Stand). This variable is named E_{WE} in EC-Lab®.

• The voltage between graphite and the reference (socket 2S and R of the PAT-Single-Stand). This variable is named E_{CE} in EC-Lab®.

• The voltage between the positive and the negative electrode. This variable is named E_{CELL} in EC-Lab®.
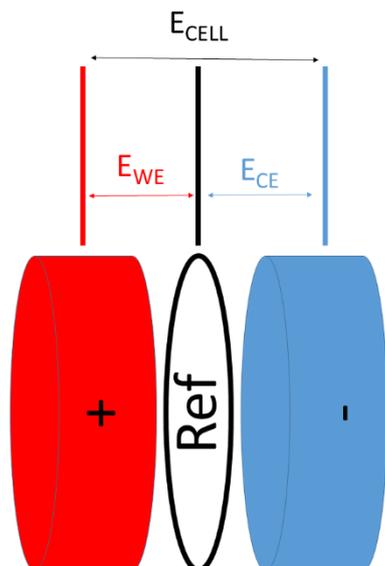
**Figure 2: Configuration for a battery with a reference electrode.**

NOTE:
During the CV period, the voltage of the cell is controlled between the positive and the negative (and not between the positive electrode and the reference as done in standard potentiostat mode).

The GCPL6 technique is available in SP-50, SP-150, VSP, VMP3 and MPG-2 instruments. Because of this specific regulation mode during the CV step, the GCPL6 cannot be linked with potentio technique. To perform an EIS measurement, it should be linked to GEIS technique.
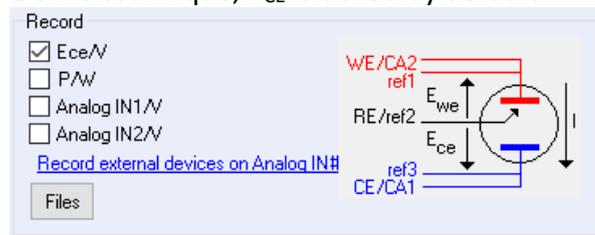
It is possible to record the two half-cell voltages and the voltage of the full cell also with SP-200, SP-240, SP-300, VSP-300 and VMP-300 instruments.

Figure 3 shows a screenshot of the GCPL6 settings, the second of overall four sequences.
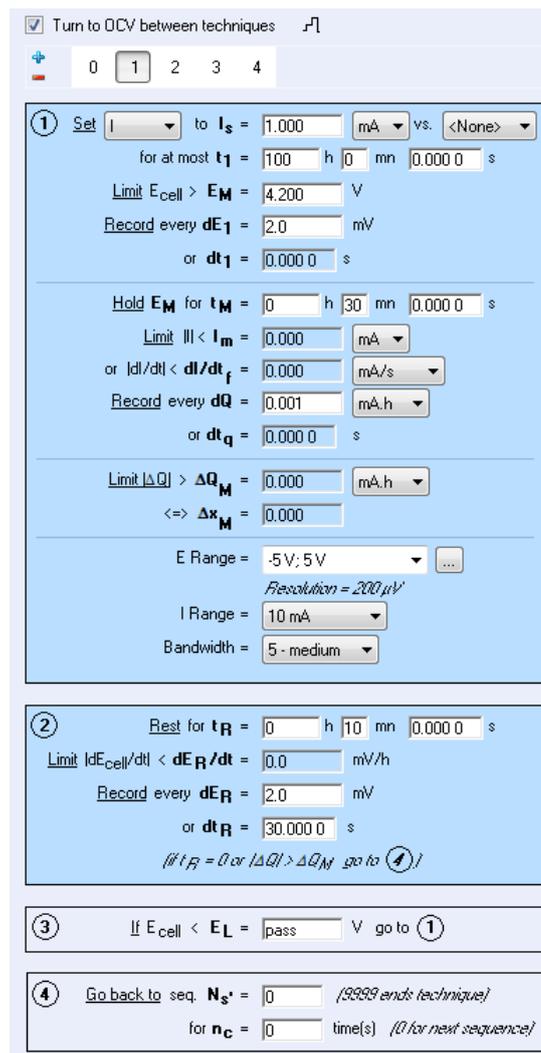
Figure 3: Sequence 1 of the GCPL6 setting.

Note that, in the Advanced Settings tab of the GCPL6 technique, E_{CE} is ticked by default.


Figure 4: E_{CE} ticked in the Advanced Setting tab.

III – CYCLING DISCUSSION

Figure 5 shows the voltage and current profiles of the overall experiment. The cut-off cell voltages were changed from 2.5/4.2 V during the initial cycle to 2.5/4.5 V in the 2nd cycle, 1.0/4.5 V in the 3rd cycle, 0.0/4.5 V in the 4th cycle, and 2.5/4.5 V in the last cycle. Throughout the experiment, the magnitude of

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the current was set to 1 mA, corresponding to a rate of approximately 0.2 C during the initial cycle.
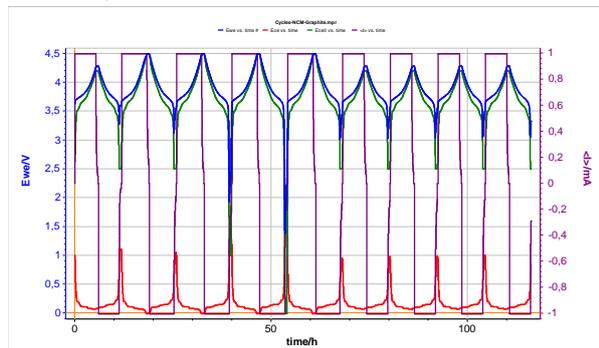
Figure 5: Voltage and current profiles of the overall experiment. The blue line corresponds to the positive half-cell voltage, the red line to the negative half-cell voltage, and the green line to the full cell voltage.

III - 1 FIRST CYCLE (BETWEEN 2.5 AND 4.2 V)

The following graphs are merely details of Fig. 5. Fig. 6 shows the evolution of the negative half-cell voltage during the formation cycle. One can clearly observe the staging plateaus of graphite. At the end of discharge, the graphite was still not fully lithiated.
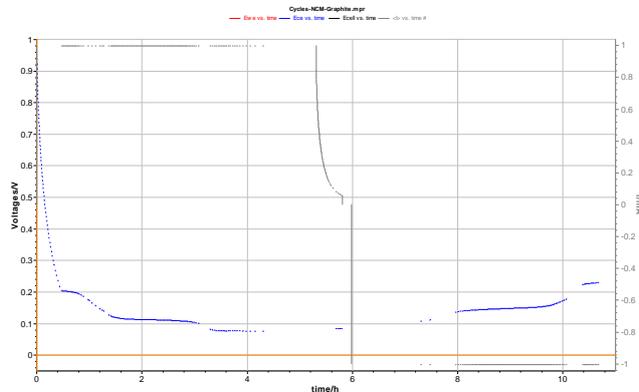

Figure 6: Evolution of the graphite half-cell voltage during first lithiation. The cut-off of the cell voltage was set to 4.2 V.

III - 2 SECOND CYCLE (BETWEEN 2.5 AND 4.5 V)

The condition changed after increasing the upper cut-off voltage to 4.5 V, see Fig. 7. When increasing the cut-off cell voltage from 4.2 to 4.5 V, the graphite electrode can no longer accommodate all the lithium released from the cathode. As a consequence, the graphite potential drops to 0 V. Plating occurs.
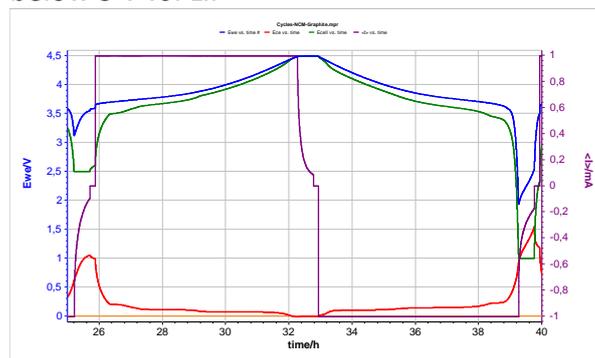
The graphite electrode is no longer able to accommodate the lithium released from the cathode. Accordingly, plating of Li metal takes place, as can be seen from the drop of the negative half-cell voltage to 0 V.
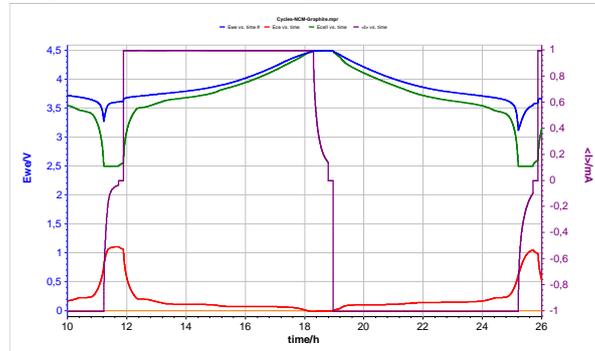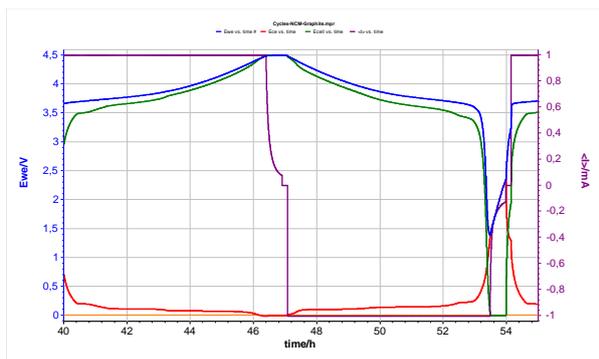

Figure 7: Voltages evolution during second cycle.

III - 3 THIRD AND FOURTH CYCLE (BETWEEN 0.0 AND 4.5 V)

During the third cycle, the lower cut-off voltage was decreased to 1.0 V, so as to see the effects of deep discharge. Discharge to below 2.5 V cell voltage is considered to potentially damage the Li-ion battery because the copper current collector of the anode may start to corrode at potentials above 3 V *vs.* Li. In the present experiment, at a cut-off voltage of 1.0 V the graphite potential does not exceed 1.6 V *vs.* Li (Fig. 8). Even when lowering the cell voltage to 0 V in the subsequent cycle (Fig. 9), the graphite potential stays well below 3 V *vs*. Li.
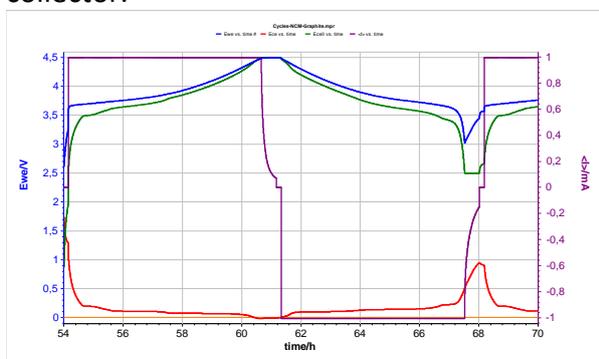

Figure 8: Voltages evolution during third cycle.

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Figure 9: Voltages evolution during the fourth cycle.

III - 4 FIFTH CYCLE (BETWEEN 2.5 AND 4.5 V)

Figure 10 depicts the last cycle of the experiment, again with the more regular cut-off cell voltages of 2.5 and 4.5 V. The battery did survive the two deep discharge cycles. The situation may change, however, when continuously cycling the battery, as the evolution of the absolute electrode potentials depends on the ratio of the two half-cell capacities. If the capacity loss of the graphite exceeds that of the NMC electrode, then the graphite electrode will continuously rise and eventually exceed the stability limit of the copper current collector.

Figure 10: Voltages evolution during the last cycle.

IV – CONCLUSION

This application note shows how to set up an experiment involving a battery with reference electrode and how to take advantage of this configuration combining an advanced test cell and cycler. This note is focused on the cycling aspect. Notably, with the same set-up described here, it is possible to measure the half and full cell impedances of the battery as well (Fig. 11). This topic will be addressed in a separate application note.
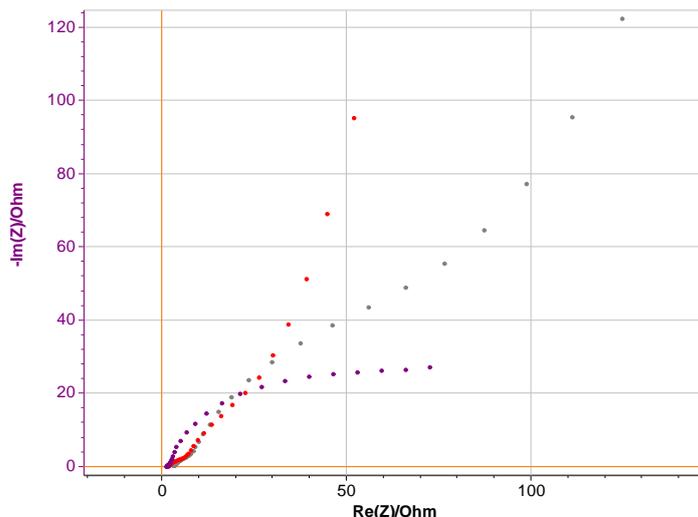
Figure 11: EIS data on NMC battery equipped with reference electrode.

Data files can be found in :
C:\Users\xxx\Documents\EC-Lab\Data\Samples\Battery\AN58 folder

REFERENCES

1) M Dolle, F. Orsini, A. S. Gozdz, and J.-M. Tarascon, J. Electrochem. Soc., 148, (2001) A851.
2) M. Klett, J. A. Gilbert, S. E. Trask, Bryant J. Polzin, A. N. Jansen, D. W. Dees, and D. P. Abraham, J. Electrochem. Soc., 163, 6 (2016) A875.
3) B. Le Gorrec, C. Montella, R. Yazami, J. Power Sources 97-97 (2001) 83.
4) More information on the cells are available on the web.
5) Technical Note #30 "Which GCPL technique is the most appropriate for my measurement?"

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